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For phosphorus,
we know that it
will make 5

bonds in the
expanded

orbital format
(Table 5.1) and
require a total
of 10

electrons,
instead of the
normal 8 that
we are used to

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seeing with the
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octet rule. So
start by

drawing in
those 5 bonds.
You'll notice
that you need
to make one
double bond
with one of the
oxygens to
complete the 5
bonds. After

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this fill in
the lone pair
electrons for
all of ...

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chemistry. This
chapter is very
important from
the point of

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your future as
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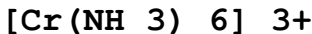
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Pdf, ... Due to
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the presence of
three unpaired

electrons in



it behaves as a
paramagnetic

substance. The

spin magnetic

moment, $\mu_s =$

$$\sqrt{g^2 S(S+1)} \mu_B =$$

$$\sqrt{3(3+2)} \mu_B =$$

$$\sqrt{15} \mu_B$$

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$\lambda = 3.87 \text{ BM}$

$[\text{Cr}(\text{NH}_3)_6]^{3+}$

is an inner

orbital

octahedral

complex. 2.

$[\text{Ni}(\text{CN})_4]^{2-}$ in

this complex Ni

is in ...

[1.3: Valence
electrons and
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Electrons In

Atoms Test

Students should be familiar with the parts of the atom from Chapter 3 but reviewing the main points is probably a good idea. Ask students questions such

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as the

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following: What
are the three
different tiny
particles that
make up an
atom? Protons,
neutrons, and
electrons.

Which of these
is in the
center of the
atom? Protons

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and neutrons
Electrons In
are in the
Atoms Test
center

(nucleus) of
the atom. You
may want to ...

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Which bond in each of the following pairs of bonds is the strongest? (a) C-C or C=C (b) C-N or C≡N (c) C≡O or C=O (d) H-F or H-Cl (e) C-H or O-H (f) C-N or C-O.

Using the bond

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energies in
Table 3,
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determine the
approximate
enthalpy change
for each of the
following
reactions: (a)
$$\text{H}_2(\text{g}) + \text{Br}_2(\text{g}) \rightarrow \dots$$

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Phosphorous in
column VA, row
3 has 5
electrons in
its outer
shell, and has
three shells in
total. Image

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from James K.

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energy | Back

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Electrons,

because they

move so fast

(approximately

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at the speed of
light), seem to
straddle the
fence

separating
energy from
matter. Albert
Einstein
developed his
famous $E=mc^2$
equation ...

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In chemistry, a group is a vertical column in the periodic table of the chemical elements. There are 18 groups in the standard periodic table,

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including the d-
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block elements
but excluding
the f-block
elements. Each
element within
a group has
similar
physical or
chemical
properties
because of its
atom's

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outermost

Electrons In

electron shell

Atoms Test

(most chemical

properties are

dominated by

the orbital

location ...

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Due to the presence of a lone pair of electrons on N, $\text{CH}_3\text{C}\equiv\text{N}$ acts as a nucleophile.

Further due to greater electronegativity of N than C, the C atom of $-\text{C}\equiv\text{N}$ carries a

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positive charge
Electrons In
and hence
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behaves as an
electrophile.

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